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Rubidium is one of the alkaline metals, as group one of the periodic table are otherwise known. The alkali metals have a single electron in their outer shell, which makes them highly reactive with oxygen, water and halogens, and also means that their oxidation state never exceeds +1.

Rubidium is a chemical element; it has symbol Rb and atomic number 37. It is a very soft, whitish-grey solid in the alkali metal group, similar to potassium and caesium. [9] Rubidium is the first alkali metal in the group to have a density higher than water.

The authors summarize the pros and cons of rubidium-82 which include low radiation exposure and a with high quality images as the pros and cost and a limited number of available PET systems being...

Rubidium atom has 37 protons while the Xenon atom has 54 protons. Protons are the positively charged particles which attract the negatively charged electrons. Now, the xenon atom has 54 protons, so it generates more attractive force compared to that of rubidium.

Rubidium is a chemical element; it has symbol Rb and atomic number 37. It is a very soft, whitish-grey solid in the alkali metal group, similar to potassium and caesium.[9] Rubidium is the first alkali metal in the group to have a density higher than water. On Earth, natural rubidium comprises two isotopes: 72% is a stable isotope 85Rb, and 28% is slightly radioactive 87Rb, with a half-life of 48.8 billion years - more than three times as long as the estimated age of the universe.

Rubidium forms a number of oxides when exposed to air, including rubidium monoxide (Rb2O), Rb6O, and Rb9O2; rubidium in excess oxygen gives the superoxide RbO2. Rubidium forms salts with halogens, producing rubidium fluoride, rubidium chloride, rubidium bromide, and rubidium iodide.[22]

Rubidium-82, one of the element"s non-natural isotopes, is produced by electron-capture decay of strontium-82 with a half-life of 25.36 days. With a half-life of 76 seconds, rubidium-82 decays by positron emission to stable krypton-82.[23]

Seawater contains an average of 125 mg/L of rubidium compared to the much higher value for potassium of 408 mg/L and the much lower value of 0.3 mg/L for caesium.[33] Rubidium is the 18th most abundant element in seawater.[15]: 371

Two notable sources of rubidium are the rich deposits of pollucite at Bernic Lake, Manitoba, Canada, and the rubicline ((Rb,K)AlSi3O8) found as impurities in pollucite on the Italian island of Elba, with a rubidium content of 17.5%.[35] Both of those deposits are also sources of caesium.[36]

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For several years in the 1950s and 1960s, a by-product of potassium production called Alkarb was a main source for rubidium. Alkarb contained 21% rubidium, with the rest being potassium and a small amount of caesium.[38] Today the largest producers of caesium produce rubidium as a by-product from pollucite.[30]

Rubidium was discovered in 1861 by Robert Bunsen and Gustav Kirchhoff, in Heidelberg, Germany, in the mineral lepidolite through flame spectroscopy. Because of the bright red lines in its emission spectrum, they chose a name derived from the Latin word rubidus, meaning "deep red".[39][40]

The slight radioactivity of rubidium was discovered in 1908, but that was before the theory of isotopes was established in 1910, and the low level of activity (half-life greater than 1010 years) made interpretation complicated. The now proven decay of 87Rb to stable 87Sr through beta decay was still under discussion in the late 1940s.[44][45]

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