



How To Easily Bring Your Dead Batteries Back To Life

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I just recently started buying batteries in 100 packs. A few short years ago, you couldn't do that. But we use so many handheld electronic devices nowadays that it's just about to the point where batteries are going to have to become a separate line item on most people's budgets.

A few people have solved this problem by switching over to rechargeables. But, rechargeable batteries are expensive, especially buying enough to operate everything that uses them. Besides, rechargeable batteries aren't going to be much help when the grid goes down and you don't have electrical power to recharge them.

The most common battery sizes - AA and AAA - provide a nominal 1.5 volts of DC power. I use the word "nominal" because that's what they are rated at. But in reality, a new AA alkaline battery actually produces slightly more power than that. As the battery is used, the voltage that it produces drops. In most cases, the device stops working when the voltage produced by the battery(s) is less than what it needs.

Few devices actually work off of one AA 1.5 volt battery. In most cases, you have to put in several. The device connects the batteries in series; which means that the positive end of one battery is connected to the negative end of the next. When batteries are connected together in this manner, the total voltage of the batteries is increased. For example:

The same applies to AAA, C and D cell batteries, which all produce 1.5 volts. The only difference is that the larger batteries can produce that amount of voltage for a longer period of time, or for devices that draw more current.

Most portable electronic devices today run off of five volts or 3.6 volts. The USB port on a computer provides five volts, so anything that can be charged off a USB charger is using a maximum of five volts. As long as the batteries in the device total less than five volts, the computer's USB port is able to charge them.

If you look at the battery of the average cell phone, you will find that it uses a lithium-Ion battery that is rated at 3.6, 3.7 or 3.8 volts, depending on the manufacturer. That's how it is able to charge off a USB charger. The voltage of the battery is less than that provided by the USB port.

You can easily find the voltage required by any electrical device by looking in the owner's manual. Even if you don't have the owner's manual, you can find out how much voltage it requires by checking the device's case where a charger connects (where it should be marked), checking the charger itself (which should also be marked) or checking the number of batteries in the device.

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At this point, the battery is normally considered "dead"; But, as I just said, the battery still has 1.3 volts of power in it. So it isn't truly dead, just partially dead. There is still power available, if we know how to access it.

Electronic devices are designed to hold the number of batteries that they need to operate. But most devices today also have a jack for connecting some sort of a charger or power adapter. So, it's possible to supply power to that device from an outside source. The device doesn't care what that outside source is, just that it provides the right amount of power.

Taking our "dead" AA batteries, we can get six volts by stringing five of them together ($5 \times 1.3\text{v} = 6.5\text{v}$); but nobody makes battery holders for five AA batteries. So, instead of stringing five of them together, we can string six together. There are battery holders for that. Six dead AA batteries will give us 7.8 volts. Since most devices have an internal voltage regulator, chances of damaging the device are virtually nonexistent.

While the battery pack can be connected directly to a fixed connector for use with a specific device, it also can be connected through a multi-connector adapter, allowing you to use various different types of connectors with it, as shown in the picture. In this case, there is a USB female adapter, which allows the typical USB to Micro USB cable for charging a phone or tablet.

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