3rd law of thermodynamics simplified

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Experimentally, it is not possible to obtain -273.15?C, as of now. It is found that most of the gases either liquify or solidify before reaching such a temperature, gaseous molecules no longer remaining. So far, scientists have been able to get close to, but not exactly, absolute zero. This may change in the future.

Frequently Asked Questions - FAQsQ1 What is thermodynamics?Thermodynamics is the branch of physical chemistry that deals with the heat, work, temperature, and energy of the system.

Entropy, denoted by "S", is a measure of the disorder or randomness in a closed system. It is directly related to the number of microstates accessible by the system, i.e. the greater the number of microstates the closed system can occupy, the greater its entropy.

The third law of thermodynamics states that the entropy of a system at absolute zero is constant or it is impossible for a process to bring the entropy of a given system to zero in a finite number of operations.

The third law of thermodynamics states that the entropy of a closed system at thermodynamic equilibrium approaches a constant value when its temperature approaches absolute zero. This constant value cannot depend on any other parameters characterizing the system, such as pressure or applied magnetic field. At absolute zero (zero kelvins) the system must be in a state with the minimum possible energy.

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